

This type of bonding is likely to take place between these types of elements:	. . . and involves this types of electron interaction :
Ionic	_____ and _____	The electrons are _____ ed between atoms
Covalent	_____ and _____	The electrons are _____ ed between atoms
Metallic	_____ and _____	The electrons are _____ ed between atoms

If electrons are shared unevenly between atoms, this is called _____ bonding.

If electrons are shared evenly between atoms, this is called _____ bonding.

X is an unknown element. If X is a nonmetal, would you expect a molecule with the formula "X₂" to have a polar bond?

Why or why not?

This type of bonding is likely to take place between these types of elements:	. . . and involves this types of electron interaction :
Ionic	<u>metal</u> and <u>nonmetal</u>	The electrons are <u>transferred</u> ed between atoms
Covalent	<u>nonmetal</u> and <u>nonmetal</u>	The electrons are <u>share</u> ed between atoms
Metallic	<u>metal</u> and <u>metal</u>	The electrons are <u>pool</u> ed between atoms

If electrons are shared unevenly between atoms, this is called polar covalent bonding.

If electrons are shared evenly between atoms, this is called nonpolar covalent bonding.

X is an unknown element. If X is a nonmetal, would you expect a molecule with the formula "X₂" to have a polar bond?

Why or why not?

No whatever's electronegativity is, the diff in electroneg. between X and X = 0, X-X → molecule of X₂

Therefore, the bond is nonpolar