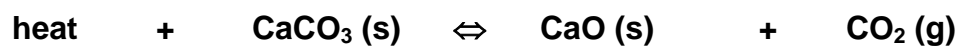


**General Chemistry**  
**Mr. MacGillivray**  
**Quiz #36:**  
**Le Châtelier's Principle**

The following reaction shows the decomposition of  $\text{CaCO}_3$  (s) at high temperatures.



How will each of these factors affect the position of equilibrium in the above equation? Answers: shift left, shift right, or NO EFFECT.

1. Addition of more  $\text{CO}_2$  (g)
2. Removal of  $\text{CO}_2$  (g)
3. Addition of  $\text{CaO}$  (s)
4. Increase in the temperature of the system
5. Increase in the pressure of the system