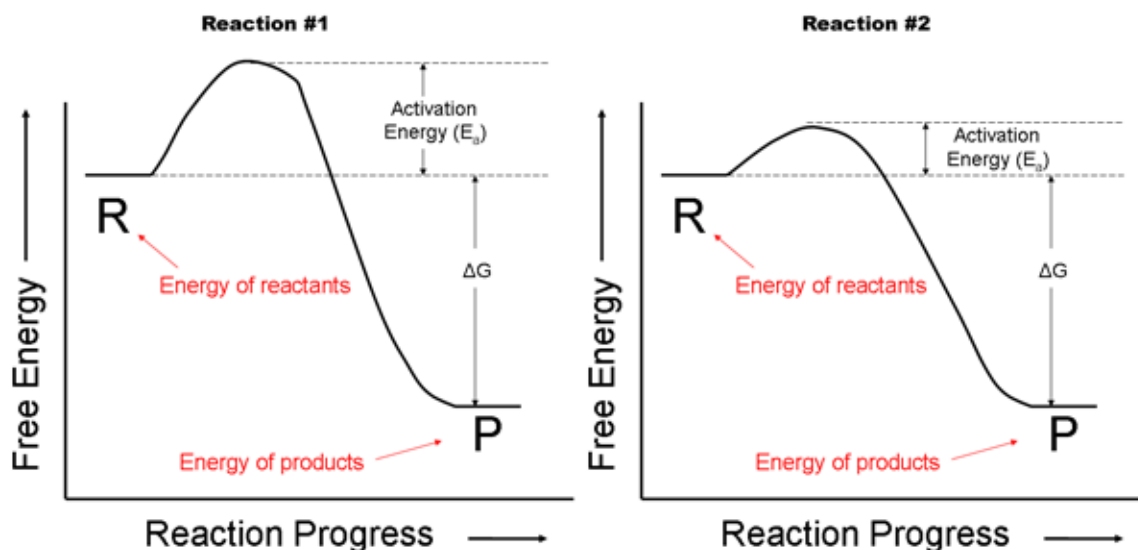


General Chemistry
Mr. MacGillivray
Quiz #34: Kinetics

How would each of the following changes affect the rate of a chemical reaction?

- | | | |
|--|----------|----------|
| 1. Adding a catalyst | INCREASE | DECREASE |
| 2. Taking one of the powdered reactants and forming it into a lump | INCREASE | DECREASE |
| 3. Increasing the concentrations of the reactants | INCREASE | DECREASE |
| 4. Decreasing the temperature | INCREASE | DECREASE |

Consider the energy diagrams below:



- | | | | |
|--|---|---|------|
| 5. Which of these reactions releases the most energy? | 1 | 2 | same |
| 6. Which reaction occurs more quickly, reaction #1 or #2? | 1 | 2 | same |
| 7. Which reaction has the highest activation energy? | 1 | 2 | same |
| 8. How will adding a catalyst affect the sign of ΔG in the reaction below ? _____ | | | |



- (a) It will change it from negative to positive
 (b) It will change it from positive to negative
 (c) It will have no effect