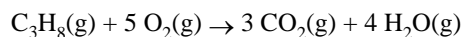


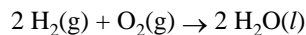
___1) Given the reaction:



At STP, what is the number of liters of CO_2 produced when 5.0 liters of $\text{C}_3\text{H}_8(\text{g})$ burns completely?

- (A) 1.0 L (C) 3.0 L
(B) 5.0 L (D) 15 L

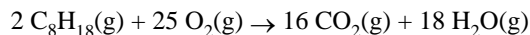
___2) Given the reaction:



What is the total number of liters of $\text{O}_2(\text{g})$ at STP needed to produce 6.0×10^{23} molecules of $\text{H}_2\text{O}(\text{l})$?

- (A) 11.2 L (C) 33.6 L
(B) 22.4 L (D) 44.8 L

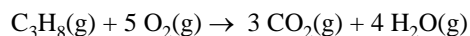
___3) Given the reaction:



What volume of $\text{C}_8\text{H}_{18}(\text{g})$ will completely react to produce exactly 36 liters of $\text{H}_2\text{O}(\text{g})$?

- (A) 27 L (C) 36 L
(B) 2.0 L (D) 4.0 L

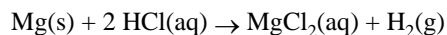
___4) Given the balanced equation:



What is the total number of liters of $\text{CO}_2(\text{g})$ produced when 20.0 liters of $\text{O}_2(\text{g})$ are completely consumed?

- (A) 12.0 L (C) 3.00 L
(B) 22.4 L (D) 5.00 L

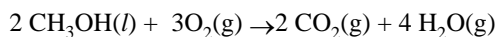
___5) Given the balanced equation:



At STP, what is the total number of liters of hydrogen gas produced when 3.00 moles of hydrochloric acid solution is completely consumed?

- (A) 11.2 L (C) 33.6 L
(B) 22.4 L (D) 44.8 L

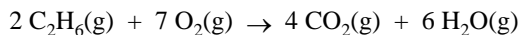
___6) Given the reaction:



How many liters of $\text{O}_2(\text{g})$ are needed to produce exactly 200 liters of $\text{CO}_2(\text{g})$?

- (A) 100 L (C) 300 L
(B) 200 L (D) 400 L

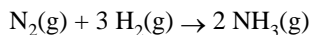
___7) Given the reaction:



At STP, what is the total volume of $\text{CO}_2(\text{g})$ formed when 6.0 liters (L) of $\text{C}_2\text{H}_6(\text{g})$ are completely oxidized?

- (A) 24 L (C) 6.0 L
(B) 12 L (D) 4.0 L

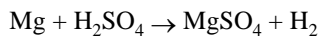
___8) Given the reaction



How many liters of ammonia, measured at STP, are produced when 28.0 grams of nitrogen is completely consumed?

- (A) 5.60 (C) 22.4
(B) 11.2 (D) 44.8

___9) Given the reaction:



How many grams of H_2SO_4 are needed to produce exactly 11.2 liters of H_2 , measured at STP?

- (A) 24.5 (C) 98.0
(B) 49.0 (D) 196

___10) Magnesium was reacted with an excess of dilute hydrochloric acid and the hydrogen gas produced collected in a eudiometer. The volume of hydrogen in the eudiometer was corrected to conditions of STP. If 94.1 milliliters of hydrogen was produced, how much magnesium reacted in this experiment?

- (A) 0.01 g (C) 0.05 g
(B) 0.10 g (D) 0.50 g