

$$\textcircled{1} 2 \text{ mol H}_2 \times \frac{4 \text{ mol PH}_3}{6 \text{ mol H}_2} = \underline{1.33 \text{ mol PH}_3} \quad \text{Hmm..... none of these answers works,}$$

$$2 \text{ mol H}_2 \times \frac{1 \text{ mol P}_4}{6 \text{ mol H}_2} = \underline{0.33 \text{ mol P}_4}$$

No answer

STOICHI  
PRACTICE  
PART I

$\textcircled{2}$  True  $\textcircled{A}$

$$\textcircled{3} 1 \text{ mol Cu} \times \frac{2 \text{ mol CuO}}{2 \text{ mol Cu}} = 1 \text{ mol CuO} \quad \textcircled{A}$$

$$\textcircled{4} 6.0 \text{ mol Br}_2 \times \frac{2 \text{ mol AlBr}_3}{3 \text{ mol Br}_2} = \underline{4} \text{ mol AlBr}_3 \quad \textcircled{C}$$

"excess Al" means there's plenty of Al to go around, so ignore it.  
"100% yield" is what we assume all of the time unless stated otherwise, so ignore that phrase, too.

$$\textcircled{5} 1.50 \text{ mol Fe} \times \frac{2 \text{ mol Fe}_2\text{O}_3}{4 \text{ mol Fe}} = \underline{0.75 \text{ mol Fe}_2\text{O}_3} \quad \textcircled{C}$$

$$\textcircled{6} \del{6.0 \text{ mol CO}_2} \times \frac{3 \text{ mol O}_2}{2 \text{ mol CO}_2} = \underline{9.0 \text{ mol O}_2} \quad \textcircled{B}$$

$$\textcircled{7} 3.0 \text{ mol KClO}_3 \times \frac{3 \text{ mol O}_2}{2 \text{ mol KClO}_3} = \underline{4.5 \text{ mol O}_2} \quad \textcircled{E}$$

$\textcircled{8}$  True  $\textcircled{A}$

$$\textcircled{9} 10.0 \text{ mol H}_2\text{O} \times \frac{4 \text{ mol NH}_3}{6 \text{ mol H}_2\text{O}} = \underline{6.67 \text{ mol NH}_3} \quad \textcircled{A}$$